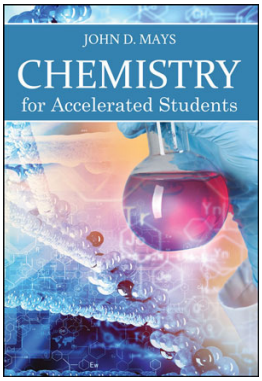
Chemistry for Accelerated Students



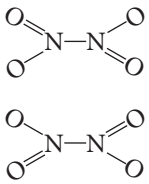
Errata

We always strive to make our textbooks as accurate as possible, but sadly, errors are a reality. We very much appreciate friends who report errata that are not included in this document!
Please send new errata to info@novarescienceandmath.com

Last revised: August 31, 2018

Chemistry for Accelerated Students 2nd ed. (2018)

Chapter 3 Exercises



12 o. N_2O_4

22a. The Be—F bond is ionic

Chemistry for Accelerated Students (2016)

Chapter 1 Exercises

16 Yb: [Xe]6s24f14
Es: [Rn]7s25f11
No: [Rn]7s25f14

Chapter 2 text

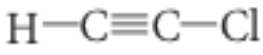
p. 52, The opening of the first paragraph should read, "The first 92 elements...are found in nature. Elements 93–118 have been synthesized in laboratories..."

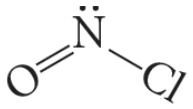
Chapter 2 Exercises

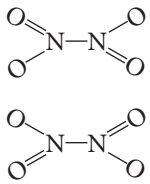
10. The problem statement should refer to cesium (Cs).

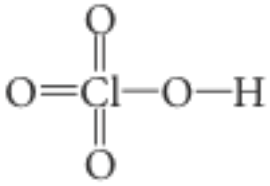
Answer: Mg < Na < Ba < Cs

Chapter 3 Answer Key

12a. 

12l. 

12 o.  N_2O_4

12s. 

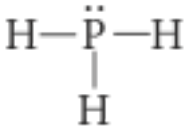
15d. perchoric acid

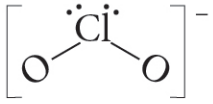
15f. bromous acid

18i. $Cr(SO_3)_3$ chromium(VI) sulfite

22a. The Be—F bond is ionic

Chapter 4 Answer Key

2a.  tetrahedral/pyramidal

2g.  tetradehral/bent

2i.  linear/linear

3e. N: *sp2* O: *sp3*

3f. sp3

Chapter 5 Answer Key

14d. $2Al(s) + 3H_2SO_4(aq) \rightarrow Al_2(SO_4)_3(aq) + 3H_2(g)$

14f. $Ba(s) + 2H_2O(l) \rightarrow Ba(OH)_2(aq) + H_2(g)$

19a.

$$750 \text{ mg } Al(OH)_3 \cdot \frac{1 \text{ g}}{1000 \text{ mg}} \cdot \frac{\text{mol}}{78.0034 \text{ g}} = 0.00961 \text{ mol } Al(OH)_3$$

$$0.00961 \text{ mol } Al(OH)_3 \cdot \frac{3 \text{ mol HCl}}{1 \text{ mol } Al(OH)_3} = 0.0288 \text{ mol HCl}$$

Rounding this result to 2 sig digs gives 0.029 mol HCl.

19b.

$$750 \text{ mg } Al(OH)_3 \cdot \frac{1 \text{ g}}{1000 \text{ mg}} \cdot \frac{\text{mol}}{78.0034 \text{ g}} = 0.00961 \text{ mol } Al(OH)_3$$

$$0.00961 \text{ mol } Al(OH)_3 \cdot \frac{3 \text{ mol } H_2O}{1 \text{ mol } Al(OH)_3} = 0.0288 \text{ mol } H_2O$$

$$0.0288 \text{ mol } H_2O \cdot \frac{18.02 \text{ g}}{\text{mol}} = 0.5198 \text{ g } H_2O$$

Rounding this result to 2 sig digs gives 0.52 g H_2O.

Chapter 9 Exercises

28 Item (g) is *basic*.

36c.

$$20.87 \text{ mL} \cdot \frac{1 \text{ L}}{1000 \text{ mL}} \cdot 4.077 \times 10^{-2} \ M = 8.509 \times 10^{-4} \text{ mol } HClO_4$$

$Ca(OH)_2$ provides two moles of OH^- ions for each mole of $Ca(OH)_2$. So only half as many moles of $Ca(OH)_2$ are required to neutralize $HClO_4$.

$$\frac{8.509 \times 10^{-4} \text{ mol } HClO_4}{2} \rightarrow 4.254 \times 10^{-4} \text{ mol } Ca(OH)_2$$

$$M_{Ca(OH)_2} \cdot 22.94 \text{ mL} \cdot \frac{1 \text{ L}}{1000 \text{ mL}} = 4.254 \times 10^{-4} \text{ mol } Ca(OH)_2$$

$$M_{Ca(OH)_2} = \frac{4.254 \times 10^{-4} \text{ mol } Ca(OH)_2}{0.02294 \text{ L}} = 0.01855 \ M \ Ca(OH)_2$$

Chapter 10 Text

Table 10-3, p. 304, Second column, fourth row should read “decreases disorder.”